2017 (S) Roll No:

# INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 2,40 Hours SUBJECTIVE MAXIMUM MARKS: 68 NOTE: - Write same question number and its part number on answer book,

as given in the question paper.

# SECTION-I

2. Attempt any eight parts.  $8 \times 2 = 16$ 

- (i) No individual neon atom in the sample of element has mass of 20.18 amu.
- Write names of any four methods employed for the separation of Isotopes. (ii)
- (iii) Differentiate between Empirical and Molecular formula.
- (iv) Define Analytical Chemistry.
- (V) Define Sublimation giving example.
- (vi) Why there is need to crystallize the crude product?
- Derive density of a gas from general gas equation (PV = nRT). (vii)
- Derive Avogadro's Law from Kinetic Theory (Kinetic equation) of gases. (viii)
- (ix) Derive units of a and b Vander Waal's Constant
- (x) What is Buffer Capacity?
- (xi) What is the effect of common ion on solubility?
- What will be the effect of change in pressure on  $NH_3$  synthesis  $N_2 + 3H_2 \implies 2NH_3$ ? (xii)

3. Attempt any eight parts.

- What is Transition Temperature? Give one example. (i)
- State Lattice Energy with an example. (ii)
- (iii) Describe Crystallographic Elements.
- (iv) State Electron Pool Theory.
- (v) State Moseley Law. Also write its importance.
- (vi) What is Azimuthal Quantum Number?
- (vii) Describe Stark and Zeeman Effect.
- (viii) Cathode rays do not depend upon the nature of the gas. Explain.
- What are Colligative Properties? Name important Colligative Properties. (ix)
- (x) NaCl Lowers the melting point of ice. Give reason.
- What is Standard Hydrogen Electrode (SHE)? (xi)
- (xii) Give some advantages of Fuel Cell.

4. Attempt any six parts.  $6 \times 2 = 12$ 

- (i) Why the distinction between Covalent Bond and Coordinate Covalent Bond vanishes after the formation of bond in  $NH_3$  and  $H_3O'$  etc.
- Why the bond angles of  $H_2O$  and  $NH_3$  are not 109.5° like that of  $CH_4$  although Oxygen (ii) and Nitrogen atoms are sp3 hybridized?
- (iii) Define Sigma and Pi bonds.

NOTE: - Attempt any three questions. 5 (a) Give any four properties of Ionic Solids.

- Define Coordinate Covalent Bond and give its one example. (iv)
- (v) What are Spontaneous and Non Spontaneous Processes? Give one example of each.
- Differentiate between Exothermic reaction and Endothermic reaction. Give one example of each. (vi)
- Justify that the order of a reaction is obtained from the rate expression of a reaction and (vii) the rate expression is obtained from the experiment.
- (viii) Why the Radioactive Decay is First Order Reaction?
- (ix) Why the rate of a chemical reaction is an ever changing parameter under the given conditions?

### SECTION-II

A sample of liquid consisting of Carbon, Hydrogen and Oxygen was subjected to combustion analysis 0.5439 g of compound gave 1.039 g of  $CO_2$  and 0.6369 g of  $H_2O_2$ . Determine empirical formula of compound. 4 Apply Kinetic Molecular Theory of gases to explain the Avogadro's Law. (b) Write the Postulates of Bohrs Atomic Model. 7.(a) Define Hybridization. Explain the geometry of  $C_2H_4$  on the basis of  $sp^2$  – Hybridization. 4 Give construction and working of Glass Colorimeter. (b)

The solubility product of  $Ag_2CrO_4$  is  $2.6 \times 10^{-2}$  at  $25^{\circ}C$ . 8.(a) Calculate the solubility of the compound.

Describe the construction and working of Standard Hydrogen Electrode. (b)

9.(a) Describe measurement of depression of freezing point by Beckmann's Method. 4 4

Define rate of reaction and explain the chemical method for its determination. (b)

4

4

4

		_	_
Pap	er	C	ode

2017 (S)

Roll No.	
----------	--

Number: 24

(B) Second order reaction

2481 INTERMEDIATE PART-I (11th CLASS)

CHEM	MISTRY PAPER-I (NEW SCHEME)	(SESS	SION 2015-2017)	
TIME		ECTIVE	MAXIMUM	MARKS: 1
think is Cutting as giver	: You have four choices for each objective type quis correct, fill that circle in front of that questioning or filling two or more circles will result in zero en in objective type question paper and leave other states are not filled. Do not solve question on this I	number. Use mark in that ers blank. No	marker or pen to fill the question. Attempt as man credit will be awarded in	circles. ny questions
(1)	The volume occupied by 1.4 g of $N_2$ at STP is: (A) 2.24 dm <sup>3</sup> (B) 22.4 dm <sup>3</sup> (C) 1.		(D) 112 cm <sup>3</sup>	
(2)	Isotopes differ in:- (A) Chemical properties		360 7-01 (1904-1970) (1	nitale
	(C) The extent to which they may be affected in a			nuis .
	(D) Properties which depend upon mass			
(3)		tion is:-		
	(A) Diethyl ether (B) Acetone (C) Et		(D) Rectified spirit	
(4)	will have the same number of molecules at	t STP.	Ta . Co	
	(A) $11.2  dm^3$ of $O_2$ and $32  g$ of $O_2$		of CO <sub>2</sub> and 280cm <sup>3</sup> of	$N_2O$
	(C) 44 g of CO <sub>2</sub> and 11.2 dm <sup>3</sup> of CO	(D) 28g of	$N_2$ and $5.6dm^3$ of $O_2$	
(5)	Pressure remaining constant at temperature what it is at 0°C? (A) 546°C (B) 22			of
(6)	Acetone and Chloroform are soluble in each oth	er due to:-	(A) Intermolecular hydrog	en bonding
	(B) dipole-dipole interaction (C) Instanta	neous dipole	(D) All of these	
(7)	Amorphous solids:- (A) Have sharp meltin	ng point		
	(B) Undergo clean cleavage when cut with knife	(C) Have pe	rfect arrangement of atoms	
	(D) Can possesses small regions of orderly arrang	gement of aton	ns	
(8)	The maximum number of electrons in an S orbit	al is:-	(A) 2 (B) 6 (C) 10	(D) 14
(9)	In the ground state of an atom, the electron is pre-	sent:-		
	(A) In the nucleus (B) In the second shell (C)	Farthesk from	n the nucleus (D) Nearest	to the nucleu
(10)	(A) $HC\ell$ (B) $HBr$ (C) $H$	IF	(D) HI	
(11)	) The geometry of $CH_4$ is:- (A) Tetrahedral	(B) Trigonal	planer (C) Linear (D) V	- Shaped
(12)	) For the reaction $NaOH + HC\ell \rightarrow NaC\ell + H_2O$	the change in	enthalpy is called:-	
	(A) Heat of reaction (B) Heat of formation (C)			lization
(13)	(A) $2NO_2 \implies N_2O_4$ (B) $N_2 + 3H_2$	$a \rightleftharpoons 2NI$	$I_3$	
(14)	(C) $H_2 + I_2 \rightleftharpoons 2HI$ (D) $2HF \rightleftharpoons$ The $pH$ of $10^{-3}$ mole $dm^{-3}$ of an aqueous solution		Fig. 1 Compared to the compare	3.0 (D) 2.7
(15)				
4-24	(B) 5.85 % solution of Sodium Chloride (C) 6.0			boiling point
(16)	F PARKET TARREST IN THE STATE OF THE STATE O		Wall	e frame
100	(A) Decrease rapidly (B) Decrease slowly (C			
(17)	C CAMP TO WATER SHAW WE SEED TO SEE THE SECOND STATE OF THE SECOND SHAW SHAW SHAW SHAW SHAW SHAW SHAW SHAW	MANUFACTURE STATE OF THE STATE	NAME OF THE PARTY	er reaction

21(Obj) (\$\frac{1}{2}\$)-2017(\$)-1300 (MULTAN)

(C) Third order reaction (D) Zero order reaction

_	_	-	_
Pa	per	Coc	ie

2017 (S)

19449-10039-1404	0404000		
Roll	No.		
100,000,000,000	06 75 Nov. 16		

# 2483 INTERMEDIATE PART-I (11th CLASS)

. vaino	3. 2105
CHEN	MISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)
ES AND S	ALLOWED: 20 Minutes OBJECTIVE MAXIMUM MARKS: 17
think is Cutting as given	You have four choices for each objective type question as A, B, C and D. The choice which you correct, fill that circle in front of that question number. Use marker or pen to fill the circles. For filling two or more circles will result in zero mark in that question. Attempt as many questions in objective type question paper and leave others blank. No credit will be awarded in case LES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.
Q.No.1	
(1)	The $pH$ of 10 <sup>-3</sup> mole $dm^{-3}$ of an aqueous solution of $H_2SO_4$ is:- (A) 2.0 (B) 1.5 (C) 3.0 (D) 2.7
(2)	solution has the highest boiling point. (A) 18.0 % solution of glucose
	(B) 5.85 % solution of Sodium Chloride (C) 6.0 % solution of Urea (D) All have same boiling points
(3)	If the salt bridge is not used between two half cells, then the voltage:-
	(A) Decrease rapidly (B) Decrease slowly (C) Drops to zero (D) Does not change
(4)	The unit of the rate constant is the same as that of the rate of reaction in:- (A) First order reaction
	(B) Second order reaction (C) Third order reaction (D) Zero order reaction
(5)	The volume occupied by 1.4 g of $N_2$ at STP is:- (A) 2.24 $dm^3$ (B) 22.4 $dm^3$ (C) 1.12 $dm^3$ (D) 112 $cm^3$
(6)	Isotopes differ in:- (A) Chemical properties (B) Arrangement of electrons in orbitals
	(C) The extent to which they may be affected in electromagnetic field
	(D) Properties which depend upon mass
(7)	The most common solvent used in solvent extraction is:-
	(A) Diethyl ether (B) Acetone (C) Ethanol (D) Rectified spirit
(8)	will have the same number of molecules at STP.
	(A) $11.2  dm^3$ of $O_2$ and $32  g$ of $O_2$ (B) $280  cm^3$ of $CO_2$ and $280  cm^3$ of $N_2O$
	(C) $44g$ of $CO_2$ and $11.2  dm^3$ of $CO$ (D) $28g$ of $N_2$ and $5.6  dm^3$ of $O_2$
(9)	Pressure remaining constant at temperature the volume of a gas will become twice of what it is at $0^{\circ}C$ ? (A) $546^{\circ}C$ (B) $220^{\circ}C$ (C) $273K$ (D) $546K$
(10)	Acctune and Chloroform are soluble in each other due to:- (A) Intermolecular hydrogen bonding
	(B) dipole-dipole interaction (C) Instantaneous dipole (D) All of these
(11)	Amorphous solids:- (A) Have sharp melting point
	(B) Undergo clean cleavage when cut with knife (C) Have perfect arrangement of atoms
	(D) Can possesses small regions of orderly arrangement of atoms
(12)	The maximum number of electrons in an $S$ orbital is:- (A) 2 (B) 6 (C) 10 (D) 14
(13)	In the ground state of an atom, the electron is present:-
	(A) In the nucleus (B) In the second shell (C) Fartherest from the nucleus (D) Nearest to the nucleus
(14)	hydrogen halides has the highest percentage of ionic character.
(15)	(A) $HC\ell$ (B) $HBr$ (C) $HF$ (D) $HI$ The geometry of $CH_4$ is:- (A) Tetrahedral (B) Trigonal planer (C) Linear (D) V – Shaped
(16)	For the reaction $NaOH + HC\ell \rightarrow NaC\ell + H_2O$ the change in enthalpy is called:-
N. L. L. L.	(A) Heat of reaction (B) Heat of formation (C) Heat of combustion (D) Heat of neutralization
(17)	For system does the equilibrium constant $K_C$ has units of (Concentration) <sup>-1</sup> .

(C)  $H_2 + I_2 \rightleftharpoons 2HI$  (D)  $2HF \rightleftharpoons H_2 + F_2$ 

(A)  $2NO_2 \implies N_2O_4$  (B)  $N_2 + 3H_2 \implies 2NH_3$ 

3					
Paper C	ode	2017	(S)	Roll No	
Numbe	2485	INTERMEDIATE	PART-I (11 <sup>th</sup>	CLASS)	
TIME A Note: Y think is Cutting as given	LLOWED: 20 : ou have four cho correct, fill that ci or filling two or m in objective type o	ices for each objective ty rele in front of that ques lore circles will result in question paper and leave	OBJECTIVE  The question as A stion number. Use zero mark in the cothers blank. N	MAXIMUM MARKS  , B, C and D. The choice which you se marker or pen to fill the circles.  It question. Attempt as many questi  No credit will be awarded in case	п
BUBBL! Q.No.1	ES are not filled.	Do not solve question of	n this sheet of Ol	SJECTIVE PAPER.	
(1)	Acetone and Chlo	oroform are soluble in eac	h other due to:-	(A) Intermolecular hydrogen bondi	ng
	(B) dipole-dipole	interaction (C) In	stantaneous dipole	(D) All of these	
(2)	Amorphous solids	s:- (A) Have sharp i	nelting point		
	(B) Undergo clean	cleavage when cut with l	knife (C) Have j	perfect arrangement of atoms	
	(D) Can possesses	small regions of orderly	arrangement of at	oms	
(3)	The maximum nur	mber of electrons in an $S$	orbital is:-	(A) 2 (B) 6 (C) 10 (D) 14	4
(4)	In the ground state	of an atom, the electron	is present:-		
	(A) In the nucleus	s (B) In the second shell	(C) Fartherest fr	om the nucleus (D) Nearest to the n	ucleus
(5)	hydrogen h (Λ) HCℓ	alides has the highest per (B) HBr	centage of ionic c (C) HF	haracter. (D) HI	
(6)	The geometry of C	$II_4$ is:- (A) Tetrah	edral (B) Trigon	al planer (C) Linear (D) V - Shape	a
(7)	For the reaction N	$aOH + HC\ell \rightarrow NaC\ell +$	$H_2O$ the change	in enthalpy is called:-	
	(A) Heat of reaction	n (B) Heat of formation	(C) Heat of cor	nbustion (D) Heat of neutralization	
(8)	(A) 2 <i>NO</i> <sub>2</sub> =	does the equilibrium cons	$+3H_2 \rightleftharpoons 2$		
(9)	The $pH$ of $10^{-3}$ n	nole dm of an aqueous	solution of $H_2SC$	O <sub>4</sub> is:- (A) 2.0 (B) 1.5 (C) 3.0 (D	)) 2.7
(10)		as the highest boiling poir			
(10)				of Urea (D) All have same boiling	points
(11)	If the salt bridge i	s not used between two h	alf cells, then the		
(12)				reaction in:- (A) First order react	ion
1	(B) Second order	reaction (C) Third	order reaction	(D) Zero order reaction	
(13)		pied by $1.4 g$ of $N_2$ at ST (B) $22.4 dm^3$		(D) 112 <i>cm</i> <sup>3</sup>	
(14)	(C) The extent to	:- (A) Chemical proper which they may be affect nich depend upon mass		rrangement of electrons in orbitals	
(15)	The most commo	on solvent used in solvent	extraction is:-		
	(A) Diethyl ethe	r (B) Acetone	(C) Ethanol	(D) Rectified spirit	Ŷ
(16)	will have	the same number of mole	cules at STP.		
	(A) 11.2 dm3 of	$O_2$ and $32 g$ of $O_2$	(B) 280 c	$m^3$ of $CO_2$ and $280  cm^3$ of $N_2O$	
	(C) 44g of C	O <sub>2</sub> and 11.2 dm <sup>3</sup> of CO	(D) 28 g	of $N_2$ and $5.6  dm^3$ of $O_2$	

(17) Pressure remaining constant at \_\_\_\_\_ temperature the volume of a gas will become twice of

what it is at  $0^{\circ}C$ ? (A)  $546^{\circ}C$ 

(B) 220°C (C) 273K

(D) 546 K

Paper Co	ode	1	2017 (S)		Roll 1	No
Number	2487	INTERMEDIA	TE PART	-I (11 <sup>th</sup> C	LASS)	
	STRY PAP LLOWED: 20	ER-I (NEW SCI Minutes	HEME) OBJEC	THE RESERVE OF THE PARTY OF THE	ION 2015-20 MAX	17) KIMUM MARKS: 17
think is ed Cutting o as given in	rrect, fill that ci filling two or n objective type	ices for each objecti ircle in front of that nore circles will resu question paper and Do not solve questi	question nur It in zero ma leave others	nber. Use rk in that blank. No	marker or pen question. Atten credit will be a	to fill the circles. npt as many questions (warded in case
(1)	will have th	ne same number of m	olecules at ST	P.		
	A) 11.2 dm <sup>3</sup> of	$O_2$ and 32 g of $O_2$	(E	3) 280 cm <sup>3</sup>	of CO <sub>2</sub> and 2	$80cm^3$ of $N_2O$
	C) 44 g of CC	$\Omega_2$ and $11.2  dm^3$ of $C$	O (E	) 28g of	$N_2$ and 5.6 dm	of O <sub>2</sub>
	Pressure remainir what it is at 0°C'	ng constant at (A) 546° C	emperature the			
(3)	Acetone and Chl (B) dipole-dipole	47			(A) Intermolecu	ılar hydrogen bonding e
(4)	Amorphous solid	s:- (A) Have sh	narp melting p	oint		
(	B) Undergo clear	n cleavage when cut v	with knife (	C) Have pe	rfect arrangeme	nt of atoms
(	D) Can possesses	s small regions of ord	erly arrangen	ent of ator	ns	
(5)	he maximum nu	mber of electrons in	an S orbital i	S:-	(A) 2 (B) 6	(C) 10 (D) 14
	SOMETHING AND AND AND	e of an atom, the elec	STATE OF THE STATE OF THE STATE OF		and the second s	
			100.00			D) Nearest to the nucleu
	(A) HCℓ	(B) HBr CH <sub>4</sub> is:- (A) T	(C) HF		(D) HI	near (D) V - Shaped
		NaOH + HCℓ → Na				
		on (B) Heat of forms				
(10)	For system	n does the equilibrium	n constant K	C has un	its of (Concentr	
	(C) $H_2 + I_2 \equiv$	<u>⇒</u> 2 <i>HI</i> (D	2 <i>HF</i> =	$H_2 + H$	72	
(11)	The <i>pH</i> of 10 <sup>-3</sup> n	nole dm <sup>-3</sup> of an aque	ous solution	of H <sub>2</sub> SO <sub>4</sub>	is:- (A) 2.0 (I	B) 1.5 (C) 3.0 (D) 2.7
(12)	solution h	as the highest boiling	point. (	A) 18.0 %	solution of gluce	ose
						have same boiling point
(13)	If the salt bridge	is not used between t	wo half cells,	then the v	oltage:-	
	(A) Decrease rapi	idly (B) Decrease	slowly (C) I	Drops to ze	ro (D) Does n	ot change
(14)	The unit of the ra	te constant is the san	e as that of th			(A) First order reaction
(	B) Second order	reaction (C)	hird order re	action	(D) Zero ord	er reaction
(15)		pied by 1.4 g of $N_2$ (B) 22.4 $dm^2$		dm¹	(D) 112cm <sup>3</sup>	
(16)	Isotopes differ in	:- (A) Chemical p	roperties	(B) Arr	angement of ele	ctrons in orbitals
	(C) The extent to	which they may be a	iffected in ele	ctromagne	tic field	
	(D) Properties wi	hich depend upon ma	SS			
(17)		on solvent used in sol			V700000000	wood where
	(A) Diethyl ethe	r (B) Acetone	(C) Etha	nol	(D) Rectified	d spirit

21(Obj) (🏗 🏗 🛣 )-2017(S)-1300 (MULTAN)

D - 11	NI-			
Roll	INO.			

# INTERMEDIATE PART-I (11th CLASS)

# CHEMISTRY PAPER-I (OLD SCHEME)

(SESSION 2012-2014)

TIME ALLOWED: 3.10 Hours

#### SUBJECTIVE

MAXIMUM MARKS: 83

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

# SECTION-I

# Attempt any eight parts.

 $8 \times 2 = 16$ 

- Define Avogadro's Number with two examples.
- (ii) Define Stioichiometric Assumptions.
- (iii) Prove  $N_2$  and CO have the same number of protons, electrons and Neutrons.
- (iv) How the crystallized substance is dried?
- (v) Define Partition Chromatography. Give example.
- (vi) Write down any two applications of Chromatography.
- (vii) What is Absolute Zero?
- (viii) Convert 80°C into F°.
- (ix) High pressure and low temperature make gas non ideal. Justify.
- (x) Define pH and pOH.
- (xi) Represent Henderson's equation for Acidic and Basic Buffer solution.
- (xii) Define Solubility.

# Attempt any eight parts.

 $8 \times 2 = 16$ 

- Boiling point of H<sub>2</sub>O is higher than that of HF. Why?
- (ii) Define Isomorphism. Give one example.
- (iii) What is a Unit Cell? Give its dimensions.
- (iv) One feels sense of cooling under the fan after bath. Give the reason.
- (v) State Pauli's Exclusion Principle.
- (vi) Write electronic configuration of  $_{29}Cu$  and  $_{24}Cr$ .
- (vii) Justify that the velocities of electrons in higher orbits are less than those in lower orbits of Hydrogen atom.
- (viii) Evaluate the Mass of Electron.
- (ix) What is Fractional Crystallization?
- (x) Differentiate between Molality and Molarity.
- (xi) How can Copper be purified Electrolytically?
- (xii) Differentiate between Electrolytic Cell and Voltaic Cell.

#### Attempt any six parts.

 $6 \times 2 = 12$ 

- (i) Second Ionization Energy is greater than first Ionization Energy. Justify it.
- (ii)  $CO_2$  has a Linear shape but  $H_2O$  has bent shape. Why?
- (iii) O, is paramagnetic in nature. Why?
- (iv) Lone pairs occupy more space than Bond Pairs. Why?
- (v) Define Spontaneous and Non-spontaneous reactions with one example each.
- (vi) State with one example, Enthalpy of Atomization.
- (vii) What is Autocatalyst? Give an example.
- (viii) What is Poisoning of a Catalyst? Give an example.
- (ix) Draw energy profile diagram of an Exothermic Reaction. Label it.

		SECTIO	N-II	
NOT	E: - Attempt	ny three questions of the following:		
5.(a) (b)	What is Limi	ing Reactant? How it is identified? don Dispersion Forces? Write the fac		on Forces. 4
6.(a) (b)	Calculate the	e sample of Hydrogen $(H_2)$ effuses for molar mass of unknown gas.		4
6.613	Explain Mills	an's oil drop experiment to determine	e/m ratio of Electron.	4
7.(a) (b)	Define Ioniza State and exp	ion Energy. Explain the factors influ- ain First Law of Thermodynamics. A	encing it. Also prove $\Delta E = q_v$	4
8.(a) (b)		racteristics of Enzyme Catalysis. oH of a buffer solution in which 0.11 r	molar $CH_3COONa$ and 0.09 m	olar acetic acid
	solutions are p	resent. Ka for CH <sub>3</sub> COOH is 1.85	× 10 <sup>-5</sup>	4
9.(a)	How is boilir by Landsberg	g point elevation of a liquid measured r's Method).	? Explain with diagram	4
(b)	What is SHE? Give diagram.	Describe a method to determine Elec-	ctrode Potential of Zinc with it.	4
		SECTION-III (PRAC	CTICAL PART)	
10.	NOTE:-	Attempt any three parts.		$(5 \times 3 = 15)$
	(i)	Write down standard solution, cher indicator with end point, procedure calculations for part A, B & C.		
	(ii)	Write down material required, dia	gram and procedure for part	D & E,
(A	) The given so	ution contains 6.3 grams of (COOH)	$(xH_2O)$ dissolved per $dm^3$ .	

- Find out value of "x" by Volumetric Method.
- (B) The given solution contains 12 grams Ammonium Oxalate  $(NH_4)_2 C_2 O_4 . H_2 O$  and Ammonium Sulphate  $(NH_4)_2 SO_4$  dissolved per  $dm^3$ . Find out the percentage of each by Volumetric Method.
- (C) The given solution contains 15 g of impure iodine dissolved per dm<sup>3</sup>. Find percentage impurity of sample by Volumetric Method.
- (D) Prepare Pure Crystals of Benzoic Acid from Impure sample by using Water as solvent.
- (E) Separate the mixture of two inks by Paper Chromatography.

Paper	Code		2017 (S)		Roll No.	
Numb	er: 6481	INTERMED	IATE PART-I	(11th CLASS)		
CHEN	MISTRY PAPI	ER-I (OLD SO				
Carrier Control of the Control	ALLOWED: 20 1		OBJECTIV			I MARKS: 17
Cutting as give	You have four chois correct, fill that cits or filling two or man in objective type of LES are not filled.	rcle in front of the ore circles will re- juestion paper and	at question numbe sult in zero mark d leave others bla	er. Use marker in that question nk. No credit w	or pen to fill tl . Attempt as m ill be awarded	e circles.
(1)	The number of mo			ygeni51-(A) 0.2	25 (B) 0.5 (C	1.0 (D) 1.5
(2)	<ul><li>(B) Is taken in less</li><li>(C) Gives the max</li></ul>	ser quantity in grar ser quantity in volu imum amount of p	ns as compared to ume as compared to roduct which is re- roduct under consi-	other reactants juired		
(3)	Solvent extraction	is an equilibrium p	process and it is cor	itrolled by:-	(A) Law of N	fass Action
	(B) The amount of	solvent used (C)	Distribution Law	(D) The amour	nt of solute	
(4)	Number of molecu					
	(A) $\frac{6.02}{22.4} \times 10^{23}$	(B) $\frac{12.04}{22.4}$	$\frac{4}{1} \times 10^{23}$ (C) $\frac{1}{2}$	$\frac{18}{2.4} \times 10^{23}$	(D) 55.6 × 6.0	$2 \times 10^{23}$
(5)	A real gas obeying	Vander Waal's eq	uation will resemb	le ideal gas if:-	(A) Both 'a' ar	nd 'b' are large
	(B) Both 'a' and 'b	are small (	<ul><li>C) 'a' is small and</li></ul>	l 'b' is large (	D) 'a' is large a	nd 'b' is small
(6)	In order to mention (A) Between 760 to (C) 765 torr	the boiling point or and 1200 torr		200 torr and 760		:-
(7)	is a pseudo : (A) $CaF_2$	solid. (B) Glass	(C) NaCl	(D) An	y type of crysta	1
(8)	The wave number	of the light emitted	l by a certain source	e is 2 x 10 <sup>6</sup> m <sup>-1</sup>		
2,802.55	The wave length of				200nm (D) 5	$\times 10^7 m$
(9)	When 6 d orbital is				100 100 0 100 100 100 100 100 100 100 1	
(3.52)	(A) 7 f	(B) 7s	(C) 7 p	(D) 7a	ſ	
(10)	The number of bon	ds in Nitrogen mo	- Evan S	12		
(*E)	(A) One σ and o			C) Three σ on	ly (D) Two σ	and one $\pi$
(11)		pole moment.		(B) <i>CHCℓ</i> ,	37 VAVAVA	
(12)	Calorie is equivale			(B) 41.84 J	S_20 W	
(13)	The pH of $10^{-3}$ mu.					-
4,000,000						
(14)	The solubility prod in the solution is:-		0 × 10 <sup>-10</sup> mol*dm <sup>-3</sup> 10 <sup>-10</sup> mol dm <sup>-3</sup>	. The maximum	concentration	of Ag* ions
	(B) $1.41 \times 10^{-5}$ mod	$dm^{-3}$ (0	(2) $1.0 \times 10^{-4}  mol  dr$	n <sup>-3</sup> (D) 4.0	$0 \times 10^{-20}  mol  dm$	-3
(15)	18g of glucose is	dissolved in 90 g o	f water. The relat	ive lowering of	Vapour pressure	is equal to:-
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{51}$	(D) 6		
(16)	If the salt bridge is	not used between	two half cells, ther	the voltage:-		
	(A) Decrease rapid	ly (B) Decrease	slowly (C) Does	not change (D)	Drops to zero	
(17)	If the rate equation	of a reaction 2A	$+ B \rightarrow \text{products is}$	s, rate = $K[A]^2$	B], and $A$ is p	resent in
	large excess, then	order of reaction is	s:- (A) 0 (	B) 1 (C) 2 (	(D) 3	
			21(Obj)( <b>X</b>	r)-2017(S)-39	O (MUL	ΓΑΝ)

CHEN	HST	RV PAD	ER-I (OLI	SCHEN	(E)	(SESS	ION 2	012-2014)	
		OWED: 20			BJECT		1011 2		M MARKS: 17
think is Cutting as given	corre or fil in ot	ect, fill that c lling two or n ojective type	ircle in front o nore circles wi	objective ty of that que ill result in r and leav	ype questi stion num zero mar e others b	on as A, ber. Use k in that lank. No	marke questio credit	will be awarde	he circles. nany questions
(1)	18g (	of glucose is	dissolved in 90	g of water.	The rela	tive lowe	ring of	Vapour pressure	is equal to:-
	(A)	$\frac{1}{5}$	(B) 5.1		(C) $\frac{1}{51}$		(D) 6	ń	
(2)	If the	e salt bridge i	s not used betw	veen two ha	alf cells, th	en the vo	ltage:-		
	(A) I	Decrease rapi	dly (B) Decre	ease slowly	(C) Do	es not cha	inge (I	) Drops to zero	
(3)	If the	e rate equation	n of a reaction	$2A + B \rightarrow$	products	is, rate =	$K[A]^2$	[B], and $A$ is $I$	present in
	larg	e excess, ther	n order of react	ion is:-	(A) 0	<b>(B)</b> 1	(C) 2	(D) 3	
(4)					a 8.0 g of	Oxygen:-	(A)	0.25 (B) 0.5 (	C) 1.0 (D) 1.5
(5)	(A) (B) (C)	Is taken in le Is taken in le Gives the ma	nt is the one wi sser quantity in sser quantity in eximum amoun nimum amount	grams as o volume as t of produc	compared t which is	to other required	reactant		
(6)	Solve	ent extraction	is an equilibriu	ım process	and it is c	ontrolled	by:-	(A) Law of M	Aass Action
	(B)	The amount of	of solvent used	(C) Distr	ibution La	w (D) T	he amo	unt of solute	
(7)			ales in one dm <sup>3</sup> (B)				10 <sup>23</sup>	(D) 55.6 × 6.	$02 \times 10^{23}$
(8)	A rea	al gas obeying	g Vander Waal	s equation	will resen	ible ideal	gas if:-	(A) Both 'a' a	nd 'b' are large
	(B)	Both 'a' and '	'b' are small	(C) 'a	' is small a	and 'b' is	large	(D) 'a' is large	and 'b' is small
(9)	(A) I		n the boiling potential torr and 1200 t	orr	er at 110°0 (B) Betwe (D) Any v	en 200 to	rr and 7	essure should be 60 torr	37 = 
(10)		is a pseud CaF <sub>2</sub>			(C) NaCl	N.	(D) A	Any type of crys	tal
(11)			er of the light en of this light wi					-1. ) 200 nm (D) :	$5 \times 10^7 m$
(12)	Wh	en 6 d orbital	is complete, th	ne entering	electron g	oes into:-			
	(A)	7 f	(B) 7s		(C) 7 p		(D)	7 d	
(13)	The	number of bo	nds in Nitroger	n molecule	is:-				
	(A)	One $\sigma$ and	one $\pi$ (B)	One σ and	l two π	(C) Th	ree σ	only (D) Two	$\sigma$ and one $\pi$
(14)	_	has zero d	ipole moment.	(	A) NH <sub>3</sub>	(B)	$CHC\ell_3$	(C) H <sub>2</sub> O	(D) BF <sub>3</sub>
(15)	Calo	rie is equival	ent to:-	(	A) 0.4184	J (B)	41.84 J	(C) 4.184 J	(D) 418.4 J
(16)	The	pH of 10 <sup>-3</sup> me	ol dm <sup>-3</sup> of an a	queous solu	ition of H	2SO4 is:-	(A) 3	.0 (B) 2.7 (C)	2.0 (D) 1.5
(17)	The	solubility pr		is 2.0 × 10	$0^{-10}  mol^2 di$			um concentratio	
						1 dm -3	(D)	$4.0 \times 10^{-20}  mol  c$	lm <sup>-3</sup>

Paper	Code		2017 (S)	RELIGIOUS PERSONNELLE PROPERTIES PARA	Roll No.
Numb	er: 6485	INTERMI	EDIATE PART	-I (11 <sup>th</sup> CLAS	S)
CHE	MISTRY PAP	ER-I (OLD	SCHEME)	(SESSION	2012-2014)
TIME	ALLOWED: 20	Minutes	OBJECT	CIVE	MAXIMUM MARKS: 1
hink is Cutting is give	s correct, fill that ci g or filling two or n n in objective type LES are not filled.	ircle in front of nore circles wil question paper	f that question num I result in zero ma and leave others	nber. Use mark rk in that quest blank. No credi	nd D. The choice which you er or pen to fill the circles. ion. Attempt as many questions t will be awarded in case VE PAPER.
(1)			rr (B) Between	C, the external een 200 torr and value of pressure	
(2)	is a pseudo (A) CaF,	solid. (B) Glass	(C) NaC	(D)	Any type of crystal
(2)	The survey assembles	nftha liabt au	15.25%-C-000000	9 100 50	,-L
(3)			itted by a certain so be:- (A) 500 nm		(a) $200 nm$ (D) $5 \times 10^7 m$
(4)	When 6 d orbital i	s complete, the	entering electron g	oes into:-	
	(A) 7 f	(B) 7s	(C) 7 p	(D)	7 <i>d</i>
(5)	The number of bon	ds in Nitrogen	molecule is:-		
	(A) One $\sigma$ and $\sigma$	one π (B) On	ne $\sigma$ and two $\pi$	(C) Three $\sigma$	only (D) Two $\sigma$ and one $\pi$
(6)	has zero dip	ole moment.	(A) NH,	(B) CHCl <sub>3</sub>	(C) $H_2O$ (D) $BF_3$
(7)	Calorie is equivaler	at to:-	(A) 0.4184	J (B) 41.84 J	(C) 4.184 J (D) 418.4 J
(8)	The pH of $10^{-3}$ mod	<i>dm</i> ⁻³ of an aqu	eous solution of H	2SO4 is:- (A)	3.0 (B) 2.7 (C) 2.0 (D) 1.5
(9)	The solubility pro in the solution is:		s $2.0 \times 10^{-10}  mol^2 d$ $0 \times 10^{-10}  mol  dm^{-3}$	$m^{-6}$ . The maxim	num concentration of $Ag^+$ ions
	(B) $1.41 \times 10^{-3}  mc$	ol dm <sup>-3</sup>	(C) $1.0 \times 10^{-4} mc$	ol dm <sup>-3</sup> (D)	$4.0 \times 10^{-20}  mol  dm^{-3}$
(10)	18 g of glucose is	dissolved in 90	g of water. The re	lative lowering of	of Vapour pressure is equal to:-
	, , 1	an et	/m 1	(D)	
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{51}$	(D)	0
(11)	If the salt bridge is	not used between	een two half cells, the	nen the voltage:-	
	(A) Decrease rapi	dly (B) Decre	ease slowly (C) D	oes not change	(D) Drops to zero
(12)	If the rate equation	on of a reaction	$2A + B \rightarrow \text{produc}$	ts is, rate = $K[A$	$\{ [B] \}$ , and $A$ is present in
	large excess, ther	n order of reacti	on is:- (A) 0	(B) 1 (C) 2	(D) 3
(13)	The number of m	oles of CO2 wb	ich contain 8.0 g of	Oxygen:- (A)	0.25 (B) 0.5 (C) 1.0 (D) 1.
(14)	A limiting reacta (A) Is taken in le (B) Is taken in le (C) Gives the ma	nt is the one whe sser quantity in sser quantity in eximum amount		I to other reactan ed to other reacta s required	ts
(15)	Solvent extractio	n is an equilibri	ium process and it i	s controlled by:-	(A) Law of Mass Action
	(B) The amount of	of solvent used	(C) Distribution I	aw (D) The an	nount of solute
(16)			of water is closed		
	(A) $\frac{6.02}{22.4} \times 10^{23}$	(B)	$\frac{12.04}{22.4} \times 10^{23} \tag{9}$	C) $\frac{18}{22.4} \times 10^{23}$	(D) $55.6 \times 6.02 \times 10^{23}$
(17)	A real gas obeying	ng Vander Waai	l's equation will res	emble ideal gas	if:- (A) Both 'a' and 'b' are larg
	(B) Both 'a' and	'b' are small	(C) 'a' is smal	and 'b' is large	(D) 'a' is large and 'b' is smal

21(Obj)(**121111**)-2017(S)-390 (MULTAN)

Paper C	code r: 6487	INTERMEDIA	2017 (S) ATE PART-I (J	1th CLASS)	Roll No				
IME A lote: Y link is lutting	ISTRY PAI ALLOWED: 20 You have four ch correct, fill that or filling two or	PER-I (OLD SCH Minutes oices for each objecti circle in front of that	HEME) (S OBJECTIV  ive type question question number  alt in zero mark i leave others blan	SESSION 201  E as A, B, C and r. Use marker n that question nk. No credit w	D. The choice which you or pen to fill the circles.  Attempt as many questions ill be awarded in case				
(1)	Number of mole (A) $\frac{6.02}{23.4} \times 10^{23}$	cules in one $dm^3$ of war (B) $\frac{12.04}{22.4}$	ater is closed to:- $\times 10^{23} \qquad (C) \frac{1}{2}$	$\frac{18}{2.4} \times 10^{23}$	(D) $55.6 \times 6.02 \times 10^{23}$				
10	Man Mar 1				(A) Both 'a' and 'b' are large				
(2)	(B) Both 'a' and		(a) 'a' is small and		(D) 'a' is large and 'b' is small				
(3)	In order to ment	ion the boiling point of torr and 1200 torr	f water at 110° C. (B) Between		essure should be:- 0 torr				
(4)	${(A)} {CaF_2}$ is a pseudon	1000 A 1000 A	(C) NaCl	(D) A	ny type of crystal				
(5)	The wave numb	per of the light emitted the of this light will be;-	by a certain source (A) 500 nm (S)	e is $2 \times 10^6 m^{-1}$ B) $500m$ (C)	200 $nm$ (D) $5 \times 10^7 m$				
(6)	When 6 d orbital is complete, the entering electron goes into:-								
		(B) 7s		(D) 7	d				
(7)		onds in Nitrogen mole	ecule is:-						
	(A) One $\sigma$ and one $\pi$ (B) One $\sigma$ and two $\pi$ (C) Three $\sigma$ only (D) Two $\sigma$ and one $\pi$								
(8)		dipole moment.	(A) NH <sub>3</sub>	(B) CHCℓ <sub>1</sub>	(C) $H_2O$ (D) $BF_3$				
(9)	Calorie is equiva		(A) 0.4184 J	(B) 41.84 J	(C) 4.184 J (D) 418.4 J				
(10)	The pH of 10	$^{3}$ mal $dm^{-3}$ of an aqueo			3.0 (B) 2.7 (C) 2.0 (D) 1.5				
(11)	The solubility in the solution	product of $AgC\ell$ is 2 is:- (A) 2.0 ×	$0 \times 10^{-10}  mol^2  dm^2$ $10^{-10}  mol  dm^{-3}$	-6. The maximu	um concentration of Ag* ions				
	(B) $1.41 \times 10^{-3}$				$4.0 \times 10^{-26}  mol  dm^{-3}$				
(12)	18 g of glucose	e is dissolved in 90 g c	of water. The rela	tive lowering of	Vapour pressure is equal to:-				
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{51}$	(D)	6				
(13	) If the salt brid	ge is not used between	n two half cells, th	en the voltage:-					
	(A) Decrease	rapidly (B) Decrease	e slowly (C) Doo	es not change (	D) Drops to zero				
(14	) If the rate equ	ation of a reaction 22	$A + B \rightarrow \text{products}$	is, rate = $K[A]$	$]^{2}[B]$ , and $A$ is present in				
	large excess.	then order of reaction	is:- (A) 0	(B) 1 (C) 2	(D) 3				
(15	The number	of moles of CO, which	h contain 8.0 g of	Oxygen:- (A)	0.25 (B) 0.5 (C) 1.0 (D)				
(16	(A) Is taken (B) Is taken (C) Gives the (D) Gives the	actant is the one which in lesser quantity in gr in lesser quantity in vo e maximum amount of e minimum amount of	n:- rams as compared blume as compared f product which is f product under co	to other reactant d to other reacta required nsideration	ts nts				
(1)	7) Solvent extra	action is an equilibriur	n process and it is	controlled by:-	(A) Law of Mass Action				
(3)	(B) The amo	unt of solvent used (	C) Distribution La	aw (D) The an	nount of solute				

# BOARD OF INTERMEDIATE AND SECONDARY EDUCATION, MULTAN OBJECTIVE KEY FOR INTER (PART-1/#) Annual Examination, 2017. The of Subject Character Session (2015-17) New

Paper Code 3481 A D C D B A B A	Paper Code 6483 CDBADCDB	Paper Code 6485  A  B  A  C  B  C	Paper Code 6487 D B A B		Rroup Q. Nos. 1. 2. 3. 4. 5.	Paper Code 2481 C D A B	Paper Code 2483 D. B C	Paper Code 2485 A D A D	Paper Code 2487 B D A
A D C D B A B A C	C D B A D C D	A B A C B D	D B A B		2. 3. 4. 5.	O D A B D	D. B. C. D. C.	A D A D	B D A D
D C D B A B A C	D B A D C D	8 A C B D	B A B		2. 3. 4. 5.	D A B D	В С D С	D A D	D A D
C B A B A	B A D C D	A C B D	A B A		3. 4. 5.	A B D	C D C	A D	A D
C B A B A	A D C D	С В D	B		4. 5.	B	D C	D	D
B B A	D C D	BD	A		5.	D	C		
A B A C	C D	D						C	Δ
A B A C	D		C	3	Ó.	0	425	the same of the sa	11_
A C		-				Α	D	A	D
A C	В	1000	В		7.	D	Α	D	C
C		В	D		8.	A	В	A	A
777	A	B	<b>C</b>		9.	D	D	D	D
В	B	C	В		10.	C	A	В	A
D	A	D	B		11.	Α	D	C	D
-		В	C		12.	D	A		В
B			D		13.	A			C
-	10000	12/14/2			14,	D		7545	D
	75.0		200	Ť	15,	В		0.00	C
					16.	C		1000	D
					17.	D	120	Total Control of	A
Ο.			>	Ť	18.				Ż
					19.				
	/				20.	- 3			
						ر مریفیکیپ سرید		s.s	<b>.</b>
S	اليدير چدهم وسلميس (د	B B D C. C D B B. B	B B D D  C: C C  D B D  B: B B  Key  المراد	B B D D B C. C C A D B D D B B D D B B B C  Key گارگر افزار	B B D D B  C. C C A  D B D D  B. B B C  Key كام اركك المركك المر	B. B A D B  13.  B D D B  14.  C. C C A  D B D D  16.  D B D D  17.  18.  19.  20.  Key على المراكز المن المن المن المن المن المن المن المن	B. B A D  B. D B  C. C C A  D B D D  13. A  14. D  15. B  D B D D  16. C  17. D  18.  19.  20.  Key عمر المراب ال	B: B A D  B D D B  C: C C A  D B D D  13. A D  14. D C  15. B A  D B D D  16. C D  17. D A  18.  19.  20.  Key عبر من عبر من عبر المراب المن المن المراب المن عبر المراب المن المراب المن المن المراب المن المن المراب المن المن المن المن المن المن المن المن	B. B A D C  B D D B  C. C C A  D B D D  15. B A A  D C  16. C D B  B. B B C  17. D A D  18.  19.  20.  Key گنا کی جو طریس کے اس طابق اور کا کا کا کا ہے۔ اس موالیہ پر چرا اور کا کا کا ہے۔ اس موالیہ پر چرا کا در اور کا در کا در اور کا در کا در اور کا در کا