2017 (S) Roll No: _____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 2.40 Hours **SUBJECTIVE** MAXIMUM MARKS: 68

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. **Attempt any eight parts.** **8 × 2 = 16**

(i) No individual neon atom in the sample of element has mass of 20.18 amu.
(ii) Write names of any four methods employed for the separation of Isotopes.
(iii) Differentiate between Empirical and Molecular formula.
(iv) Define Analytical Chemistry.
(v) Define Sublimation giving example.
(vi) Why there is need to crystallize the crude product?
(vii) Derive density of a gas from general gas equation ($PV = nRT$).
(viii) Derive Avogadro's Law from Kinetic Theory (Kinetic equation) of gases.
(ix) Derive units of a and b Vander Waal's Constant
(x) What is Buffer Capacity?
(xi) What is the effect of common ion on solubility?
(xii) What will be the effect of change in pressure on NH_3 synthesis $N_2 + 3H_2 \rightleftharpoons 2NH_3$?

3. **Attempt any eight parts.** **8 × 2 = 16**

(i) What is Transition Temperature? Give one example.
(ii) State Lattice Energy with an example.
(iii) Describe Crystallographic Elements.
(iv) State Electron Pool Theory.
(v) State Moseley Law. Also write its importance.
(vi) What is Azimuthal Quantum Number?
(vii) Describe Stark and Zeeman Effect.
(viii) Cathode rays do not depend upon the nature of the gas. Explain.
(ix) What are Colligative Properties? Name important Colligative Properties.
(x) $NaCl$ Lowers the melting point of ice. Give reason.
(xi) What is Standard Hydrogen Electrode (SHE)?
(xii) Give some advantages of Fuel Cell.

4. **Attempt any six parts.** **6 × 2 = 12**

(i) Why the distinction between Covalent Bond and Coordinate Covalent Bond vanishes after the formation of bond in NH_3 and H_3O^+ etc.
(ii) Why the bond angles of H_2O and NH_3 are not 109.5° like that of CH_4 although Oxygen and Nitrogen atoms are sp^3 hybridized?
(iii) Define Sigma and Pi bonds.
(iv) Define Coordinate Covalent Bond and give its one example.
(v) What are Spontaneous and Non Spontaneous Processes? Give one example of each.
(vi) Differentiate between Exothermic reaction and Endothermic reaction. Give one example of each.
(vii) Justify that the order of a reaction is obtained from the rate expression of a reaction and the rate expression is obtained from the experiment.
(viii) Why the Radioactive Decay is First Order Reaction?
(ix) Why the rate of a chemical reaction is an ever changing parameter under the given conditions?

SECTION-II

NOTE: - Attempt any three questions.

5.(a) Give any four properties of Ionic Solids. 4
(b) A sample of liquid consisting of Carbon, Hydrogen and Oxygen was subjected to combustion analysis 0.5439 g of compound gave 1.039 g of CO_2 and 0.6369 g of H_2O. Determine empirical formula of compound. 4

6.(a) Apply Kinetic Molecular Theory of gases to explain the Avogadro's Law. 4
(b) Write the Postulates of Bohrs Atomic Model. 4

7.(a) Define Hybridization. Explain the geometry of C_2H_4 on the basis of sp^2 – Hybridization. 4
(b) Give construction and working of Glass Colorimeter. 4

8.(a) The solubility product of Ag_2CrO_4 is 2.6×10^{-4} at $25^\circ C$. Calculate the solubility of the compound. 4
(b) Describe the construction and working of Standard Hydrogen Electrode. 4

9.(a) Describe measurement of depression of freezing point by Beckmann's Method. 4
(b) Define rate of reaction and explain the chemical method for its determination. 4

Paper Code Number: 2481

2017 (S) Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

(1) The volume occupied by 1.4 g of N_2 at STP is:-
(A) 2.24 dm^3 (B) 22.4 dm^3 (C) 1.12 dm^3 (D) 112 cm^3

(2) Isotopes differ in:- (A) Chemical properties (B) Arrangement of electrons in orbitals
(C) The extent to which they may be affected in electromagnetic field
(D) Properties which depend upon mass

(3) The most common solvent used in solvent extraction is:-
(A) Diethyl ether (B) Acetone (C) Ethanol (D) Rectified spirit

(4) _____ will have the same number of molecules at STP.
(A) 11.2 dm^3 of O_2 and 32 g of O_2 (B) 280 cm^3 of CO_2 and 280 cm^3 of N_2O
(C) 44 g of CO_2 and 11.2 dm^3 of CO (D) 28 g of N_2 and 5.6 dm^3 of O_2

(5) Pressure remaining constant at _____ temperature the volume of a gas will become twice of what it is at 0° C? (A) 546° C (B) 220° C (C) 273 K (D) 546 K

(6) Acetone and Chloroform are soluble in each other due to:- (A) Intermolecular hydrogen bonding
(B) dipole-dipole interaction (C) Instantaneous dipole (D) All of these

(7) Amorphous solids:- (A) Have sharp melting point
(B) Undergo clean cleavage when cut with knife (C) Have perfect arrangement of atoms
(D) Can possesses small regions of orderly arrangement of atoms

(8) The maximum number of electrons in an S orbital is:- (A) 2 (B) 6 (C) 10 (D) 14

(9) In the ground state of an atom, the electron is present:-
(A) In the nucleus (B) In the second shell (C) Farthest from the nucleus (D) Nearest to the nucleus

(10) _____ hydrogen halides has the highest percentage of ionic character.
(A) HCl (B) HBr (C) HF (D) HI

(11) The geometry of CH_4 is:- (A) Tetrahedral (B) Trigonal planer (C) Linear (D) V – Shaped

(12) For the reaction $NaOH + HCl \rightarrow NaCl + H_2O$ the change in enthalpy is called:-
(A) Heat of reaction (B) Heat of formation (C) Heat of combustion (D) Heat of neutralization

(13) For _____ system does the equilibrium constant K_C has units of $(Concentration)^{-1}$.
(A) $2NO_2 \rightleftharpoons N_2O_4$ (B) $N_2 + 3H_2 \rightleftharpoons 2NH_3$
(C) $H_2 + I_2 \rightleftharpoons 2HI$ (D) $2HF \rightleftharpoons H_2 + F_2$

(14) The pH of 10^{-3} mole dm^{-3} of an aqueous solution of H_2SO_4 is:- (A) 2.0 (B) 1.5 (C) 3.0 (D) 2.7

(15) _____ solution has the highest boiling point. (A) 18.0 % solution of glucose
(B) 5.85 % solution of Sodium Chloride (C) 6.0 % solution of Urea (D) All have same boiling points

(16) If the salt bridge is not used between two half cells, then the voltage:-
(A) Decrease rapidly (B) Decrease slowly (C) Drops to zero (D) Does not change

(17) The unit of the rate constant is the same as that of the rate of reaction in:- (A) First order reaction
(B) Second order reaction (C) Third order reaction (D) Zero order reaction

Paper Code Number: 2483

2017 (S) Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

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(2) _____ solution has the highest boiling point. (A) 18.0 % solution of glucose

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(3) If the salt bridge is not used between two half cells, then the voltage:-

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(B) Second order reaction (C) Third order reaction (D) Zero order reaction

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(D) Properties which depend upon mass

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(8) _____ will have the same number of molecules at STP.

(A) 11.2 dm^3 of O_2 and 32 g of O_2 (B) 280 cm^3 of CO_2 and 280 cm^3 of N_2O

(C) 44 g of CO_2 and 11.2 dm^3 of CO (D) 28 g of N_2 and 5.6 dm^3 of O_2

(9) Pressure remaining constant at _____ temperature the volume of a gas will become twice of what it is at 0°C? (A) 546°C (B) 220°C (C) 273K (D) 546 K

(10) Acetone and Chloroform are soluble in each other due to:- (A) Intermolecular hydrogen bonding

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(12) The maximum number of electrons in an S orbital is:- (A) 2 (B) 6 (C) 10 (D) 14

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(16) For the reaction $NaOH + HCl \rightarrow NaCl + H_2O$ the change in enthalpy is called:-

(A) Heat of reaction (B) Heat of formation (C) Heat of combustion (D) Heat of neutralization

(17) For _____ system does the equilibrium constant K_C has units of $(Concentration)^{-1}$.

(A) $2NO_2 \rightleftharpoons N_2O_4$ (B) $N_2 + 3H_2 \rightleftharpoons 2NH_3$

(C) $H_2 + I_2 \rightleftharpoons 2HI$ (D) $2HF \rightleftharpoons H_2 + F_2$

Paper Code Number: 2485	2017 (S)	Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

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(3) The maximum number of electrons in an S orbital is:- (A) 2 (B) 6 (C) 10 (D) 14

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(9) The pH of 10^{-3} mole dm^{-3} of an aqueous solution of H_2SO_4 is:- (A) 2.0 (B) 1.5 (C) 3.0 (D) 2.7

(10) _____ solution has the highest boiling point. (A) 18.0 % solution of glucose (B) 5.85 % solution of Sodium Chloride (C) 6.0 % solution of Urea (D) All have same boiling points

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(17) Pressure remaining constant at _____ temperature the volume of a gas will become twice of what it is at 0° C? (A) 546° C (B) 220° C (C) 273 K (D) 546 K

Paper Code Number: 2487	2017 (S)	Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

(1) _____ will have the same number of molecules at STP.

(A) $11.2\,dm^3$ of O_2 and $32\,g$ of O_2 (B) $280\,cm^3$ of CO_2 and $280\,cm^3$ of N_2O

(C) $44\,g$ of CO_2 and $11.2\,dm^3$ of CO (D) $28\,g$ of N_2 and $5.6\,dm^3$ of O_2

(2) Pressure remaining constant at _____ temperature the volume of a gas will become twice of what it is at $0^\circ C$? (A) $546^\circ C$ (B) $220^\circ C$ (C) $273K$ (D) $546\,K$

(3) Acetone and Chloroform are soluble in each other due to:- (A) Intermolecular hydrogen bonding (B) dipole-dipole interaction (C) Instantaneous dipole (D) All of these

(4) Amorphous solids:- (A) Have sharp melting point (B) Undergo clean cleavage when cut with knife (C) Have perfect arrangement of atoms (D) Can possesses small regions of orderly arrangement of atoms

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(16) Isotopes differ in:- (A) Chemical properties (B) Arrangement of electrons in orbitals (C) The extent to which they may be affected in electromagnetic field (D) Properties which depend upon mass

(17) The most common solvent used in solvent extraction is:- (A) Diethyl ether (B) Acetone (C) Ethanol (D) Rectified spirit

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 3.10 Hours **SUBJECTIVE** MAXIMUM MARKS: 83

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any eight parts. **8 × 2 = 16**

(i) Define Avogadro's Number with two examples.

(ii) Define Stioichiometric Assumptions.

(iii) Prove N_2 and CO have the same number of protons, electrons and Neutrons.

(iv) How the crystallized substance is dried?

(v) Define Partition Chromatography. Give example.

(vi) Write down any two applications of Chromatography.

(vii) What is Absolute Zero?

(viii) Convert $80^{\circ}C$ into F°.

(ix) High pressure and low temperature make gas non ideal. Justify.

(x) Define pH and pOH.

(xi) Represent Henderson's equation for Acidic and Basic Buffer solution.

(xii) Define Solubility.

3. Attempt any eight parts. **8 × 2 = 16**

(i) Boiling point of H_2O is higher than that of HF. Why?

(ii) Define Isomorphism. Give one example.

(iii) What is a Unit Cell? Give its dimensions.

(iv) One feels sense of cooling under the fan after bath. Give the reason.

(v) State Pauli's Exclusion Principle.

(vi) Write electronic configuration of $_{29}Cu$ and $_{24}Cr$.

(vii) Justify that the velocities of electrons in higher orbits are less than those in lower orbits of Hydrogen atom.

(viii) Evaluate the Mass of Electron.

(ix) What is Fractional Crystallization?

(x) Differentiate between Molality and Molarity.

(xi) How can Copper be purified Electrolytically?

(xii) Differentiate between Electrolytic Cell and Voltaic Cell.

4. Attempt any six parts. **6 × 2 = 12**

(i) Second Ionization Energy is greater than first Ionization Energy. Justify it.

(ii) CO_2 has a Linear shape but H_2O has bent shape. Why?

(iii) O_2 is paramagnetic in nature. Why?

(iv) Lone pairs occupy more space than Bond Pairs. Why?

(v) Define Spontaneous and Non-spontaneous reactions with one example each.

(vi) State with one example, Enthalpy of Atomization.

(vii) What is Autocatalyst? Give an example.

(viii) What is Poisoning of a Catalyst? Give an example.

(ix) Draw energy profile diagram of an Exothermic Reaction. Label it.

SECTION-II

NOTE: - Attempt any three questions of the following:-

5.(a) What is Limiting Reactant? How it is identified? 4

(b) What are London Dispersion Forces? Write the factors affecting London Dispersion Forces. 4

6.(a) $250\,cm^3$ of the sample of Hydrogen (H_2) effuses four times as rapidly as $250\,cm^3$ of an unknown gas. Calculate the molar mass of unknown gas. 4

(b) Explain Milikan's oil drop experiment to determine e/m ratio of Electron. 4

7.(a) Define Ionization Energy. Explain the factors influencing it. 4

(b) State and explain First Law of Thermodynamics. Also prove $\Delta E = q_v$ 4

8.(a) Give four characteristics of Enzyme Catalysis. 4

(b) Calculate the pH of a buffer solution in which 0.11 molar CH_3COONa and 0.09 molar acetic acid solutions are present. Ka for CH_3COOH is 1.85×10^{-5} 4

9.(a) How is boiling point elevation of a liquid measured? Explain with diagram by Landsberger's Method). 4

(b) What is SHE? Describe a method to determine Electrode Potential of Zinc with it. Give diagram. 4

SECTION-III (PRACTICAL PART)

10. NOTE:- Attempt any three parts. (5 x 3 = 15)

(i) Write down standard solution, chemical equation with mole ratio, indicator with end point, procedure and supposed readings with calculations for part A, B & C.

(ii) Write down material required, diagram and procedure for part D & E.

(A) The given solution contains 6.3 grams of $(COOH)_2.xH_2O$ dissolved per dm^3. Find out value of "x" by Volumetric Method.

(B) The given solution contains 12 grams Ammonium Oxalate $(NH_4)_2C_2O_4.H_2O$ and Ammonium Sulphate $(NH_4)_2SO_4$ dissolved per dm^3. Find out the percentage of each by Volumetric Method.

(C) The given solution contains 15 g of impure iodine dissolved per dm^3. Find percentage impurity of sample by Volumetric Method.

(D) Prepare Pure Crystals of Benzoic Acid from Impure sample by using Water as solvent.

(E) Separate the mixture of two inks by Paper Chromatography.

Paper Code Number: 6481	2017 (S) INTERMEDIATE PART-I (11th CLASS)	Roll No._____

CHEMISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

(1) The number of moles of CO_2 which contain 8.0 g of Oxygen is:- (A) 0.25 (B) 0.5 (C) 1.0 (D) 1.5

(2) A limiting reactant is the one which:-
(A) Is taken in lesser quantity in grams as compared to other reactants
(B) Is taken in lesser quantity in volume as compared to other reactants
(C) Gives the maximum amount of product which is required
(D) Gives the minimum amount of product under consideration

(3) Solvent extraction is an equilibrium process and it is controlled by:- (A) Law of Mass Action
(B) The amount of solvent used (C) Distribution Law (D) The amount of solute

(4) Number of molecules in one dm^3 of water is closed to:-
(A) $\frac{6.02}{22.4} \times 10^{23}$ (B) $\frac{12.04}{22.4} \times 10^{23}$ (C) $\frac{18}{22.4} \times 10^{23}$ (D) $55.6 \times 6.02 \times 10^{23}$

(5) A real gas obeying Vander Waal's equation will resemble ideal gas if:- (A) Both 'a' and 'b' are large
(B) Both 'a' and 'b' are small (C) 'a' is small and 'b' is large (D) 'a' is large and 'b' is small

(6) In order to mention the boiling point of water at $110^{\circ}C$, the external pressure should be:-
(A) Between 760 torr and 1200 torr (B) Between 200 torr and 760 torr
(C) 765 torr (D) Any value of pressure

(7) _____ is a pseudo solid.
(A) CaF_2 (B) Glass (C) $NaC\ell$ (D) Any type of crystal

(8) The wave number of the light emitted by a certain source is $2 \times 10^6 m^{-1}$.
The wave length of this light will be:- (A) $500\,nm$ (B) $500\,m$ (C) $200\,nm$ (D) $5 \times 10^7 m$

(9) When 6 d orbital is complete, the entering electron goes into:-
(A) $7f$ (B) $7s$ (C) $7p$ (D) $7d$

(10) The number of bonds in Nitrogen molecule is:-
(A) One σ and one π (B) One σ and two π (C) Three σ only (D) Two σ and one π

(11) _____ has zero dipole moment. (A) NH_3 (B) $CHC\ell_3$ (C) H_2O (D) BF_3

(12) Calorie is equivalent to:- (A) 0.4184 J (B) 41.84 J (C) 4.184 J (D) 418.4 J

(13) The pH of $10^{-3}\,mol\,dm^{-3}$ of an aqueous solution of H_2SO_4 is:- (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5

(14) The solubility product of $AgC\ell$ is $2.0 \times 10^{-10}\,mol^2 dm^{-6}$. The maximum concentration of Ag^+ ions in the solution is:- (A) $2.0 \times 10^{-10}\,mol\,dm^{-3}$
(B) $1.41 \times 10^{-5}\,mol\,dm^{-3}$ (C) $1.0 \times 10^{-4}\,mol\,dm^{-3}$ (D) $4.0 \times 10^{-20}\,mol\,dm^{-3}$

(15) $18\,g$ of glucose is dissolved in 90 g of water. The relative lowering of Vapour pressure is equal to:-
(A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6

(16) If the salt bridge is not used between two half cells, then the voltage:-
(A) Decrease rapidly (B) Decrease slowly (C) Does not change (D) Drops to zero

(17) If the rate equation of a reaction $2A + B \rightarrow$ products is, rate $= K[A]^2[B]$, and A is present in large excess, then order of reaction is:- (A) 0 (B) 1 (C) 2 (D) 3

Paper Code Number: 6483

2017 (S) Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

(1) $18\,g$ of glucose is dissolved in 90 g of water. The relative lowering of Vapour pressure is equal to:-

(A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6

(2) If the salt bridge is not used between two half cells, then the voltage:-

(A) Decrease rapidly (B) Decrease slowly (C) Does not change (D) Drops to zero

(3) If the rate equation of a reaction $2A + B \rightarrow$ products is, rate = $K[A]^2[B]$, and A is present in large excess, then order of reaction is:- (A) 0 (B) 1 (C) 2 (D) 3

(4) The number of moles of CO_2 which contain 8.0 g of Oxygen:- (A) 0.25 (B) 0.5 (C) 1.0 (D) 1.5

(5) A limiting reactant is the one which:-
(A) Is taken in lesser quantity in grams as compared to other reactants
(B) Is taken in lesser quantity in volume as compared to other reactants
(C) Gives the maximum amount of product which is required
(D) Gives the minimum amount of product under consideration

(6) Solvent extraction is an equilibrium process and it is controlled by:- (A) Law of Mass Action (B) The amount of solvent used (C) Distribution Law (D) The amount of solute

(7) Number of molecules in one dm^3 of water is closed to:-

(A) $\frac{6.02}{22.4} \times 10^{23}$ (B) $\frac{12.04}{22.4} \times 10^{23}$ (C) $\frac{18}{22.4} \times 10^{23}$ (D) $55.6 \times 6.02 \times 10^{23}$

(8) A real gas obeying Vander Waal's equation will resemble ideal gas if:- (A) Both 'a' and 'b' are large (B) Both 'a' and 'b' are small (C) 'a' is small and 'b' is large (D) 'a' is large and 'b' is small

(9) In order to mention the boiling point of water at $110^\circ C$, the external pressure should be:-
(A) Between 760 torr and 1200 torr (B) Between 200 torr and 760 torr
(C) 765 torr (D) Any value of pressure

(10) _____ is a pseudo solid.
(A) CaF_2 (B) Glass (C) $NaC\ell$ (D) Any type of crystal

(11) The wave number of the light emitted by a certain source is $2 \times 10^6\,m^{-1}$.
The wave length of this light will be:- (A) $500\,nm$ (B) $500\,m$ (C) $200\,nm$ (D) $5 \times 10^7\,m$

(12) When 6 d orbital is complete, the entering electron goes into:-

(A) $7f$ (B) $7s$ (C) $7p$ (D) $7d$

(13) The number of bonds in Nitrogen molecule is:-

(A) One σ and one π (B) One σ and two π (C) Three σ only (D) Two σ and one π

(14) _____ has zero dipole moment. (A) NH_3 (B) $CHC\ell_3$ (C) H_2O (D) BF_3

(15) Calorie is equivalent to:- (A) 0.4184 J (B) 41.84 J (C) 4.184 J (D) 418.4 J

(16) The pH of $10^{-3}\,mol\,dm^{-3}$ of an aqueous solution of H_2SO_4 is:- (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5

(17) The solubility product of $AgC\ell$ is $2.0 \times 10^{-10}\,mol^2\,dm^{-6}$. The maximum concentration of Ag^+ ions in the solution is:- (A) $2.0 \times 10^{-10}\,mol\,dm^{-3}$
(B) $1.41 \times 10^{-5}\,mol\,dm^{-3}$ (C) $1.0 \times 10^{-4}\,mol\,dm^{-3}$ (D) $4.0 \times 10^{-20}\,mol\,dm^{-3}$

Paper Code Number: 6485	2017 (S) INTERMEDIATE PART-I (11th CLASS)	Roll No._____

CHEMISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014)

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Q.No.1

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Paper Code Number: 6487	2017 (S)	Roll No._____

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014)

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BOARD OF INTERMEDIATE AND SECONDARY EDUCATION, MULTAN

OBJECTIVE KEY FOR INTER (PART-I / II) Annual Examination, 2017.

(اولڈ سکیم) Name of Subject Chemistry — Group: I (2012 – 14)

Session (2015-17) (New Scheme) — Group: 2nd

Q. Nos.	Paper Code 6481	Paper Code 6483	Paper Code 6485	Paper Code 6487
1.	A	C	A	D
2.	D	D	B	B
3.	C	B	A	A
4.	D	A	C	B
5.	B	D	B	A
6.	A	C	D	C
7.	B	D	C	B
8.	A	B	B	D
9.	C	A	B	C
10.	B	B	C	B
11.	D	A	D	B
12.	C	C	B	C
13.	B	B	A	D
14.	B	D	D	B
15.	C	C	C	A
16.	D	B	D	D
17.	B	B	B	C
18.				
19.				
20.				

Q. Nos.	Paper Code 2481	Paper Code 2483	Paper Code 2485	Paper Code 2487
1.	C	D	A	B
2.	D	B	D	D
3.	A	C	A	A
4.	B	D	D	D
5.	D	C	C	A
6.	A	D	A	D
7.	D	A	D	C
8.	A	B	A	A
9.	D	D	D	D
10.	C	A	B	A
11.	A	D	C	D
12.	D	A	D	B
13.	A	D	C	C
14.	D	C	D	D
15.	B	A	A	C
16.	C	D	B	D
17.	D	A	D	A
18.				
19.				
20.				

سرٹیفکیٹ بابت تصحیح سوالیہ پرچہ / مارکنگ Key

ہم نے مضمون کیمسٹری پرچہ پارٹ I گروپ I میڈیم اردو / انگلش انٹر سالانہ / ضمنی امتحان 2017 کا سوالیہ پرچہ / نیا معروضی (Subjective & Objevtive) کو بنظر عمیق چیک کرلیا ہے یہ پرچہ سلیبس کے عین مطابق Set کیا گیا ہے۔ اس سوالیہ پرچہ میں کسی قسم کی کوئی غلطی نہ ہے۔ ہم نے سوالیہ پرچہ کا اردو اور انگریزی Version بھی چیک کرلیا ہے یہ Version آپس میں مطابقت رکھتے ہیں اور سلیبس (Syllabus) کے مطابق بھی ہیں۔ نیز اس پرچہ کی Key کی بابت بھی تصدیق کی جاتی ہے کہ یہ بھی درست بنائی گئی ہے۔ اس میں بھی کسی قسم کی کوئی غلطی نہ ہے۔
مزید یہ کہ ہم نے Key بنانے سے متعلق دفتر کی جانب سے تیار کردہ ہدایات وصول کر کے ان کا بغور مطالعہ کرلیا ہے اور ان کی روشنی میں Key بنائی ہے۔

PREPARED & CHECKED BY

Sr.No	Name	Designation	Institution	Mobile No.	Signature.
1	Dr. M. Ramzan	A.P	GEC. Multan	03346027542	[signature]
2	Javed Akhtar	Associate	Science College Multan	03007336924	[signature]
3	M. Khalid Bhatti	Associate (R) Prof.	GC Civil Lines Multan	03339261304	[signature]
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