2017 (S) Roll No: _____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 2.40 Hours **SUBJECTIVE** MAXIMUM MARKS: 68

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any eight parts. **8 × 2 = 16**

(i) Diamond is electrically insulator but graphite is fairly good conductor. Justify it.
(ii) Justify that Ionization Energy value decreases down the group.
(iii) Write advantages of DownsCell for preparation of Sodium (Na).
(iv) Justify that CO_2 in excess when passed through Lime Water solution becomes clear. Write reaction involved.
(v) Write down formula of (i) Cryolite (ii) Colemanite
(vi) How will you explain structure of Carbon Monoxide (CO)?
(vii) Write effect of temperature on N_2O_4. Give colour change and reaction.
(viii) Justify that Aqua Regia can dissolve nearly all types of substances.
(ix) What is the difference between Typical and Non-typical Transition elements?
(x) How Steel is classified on the base of Carbon contents?
(xi) Write only name of various sources which contaminate surface and ground water.
(xii) Write importance of gases present in the atmosphere.

3. Attempt any eight parts. **8 × 2 = 16**

(i) Why HF is weaker acid than HCl?
(ii) What is Iodized Salt?
(iii) Why there is no free rotation around a double bond and a free rotation around a single bond?
(iv) Define Functional Group Isomerism. Give one example.
(v) Alkanes are less reactive than Alkenes. Why?
(vi) Ethene can be converted into Ethyl Alcohol. Write equation.
(vii) Draw structural formulas for:- (i) p – Nitroaniline (ii) m – Nitrophenol
(viii) What happens when Chlorine is passed through Benzene in Sunlight?
(ix) Define Nucleophilic substitution reactions. Giver their general mechanism.
(x) Define Alkyl Halides. Which is the best method of preparing Alkyl Halides?
(xi) Write the structural formulas of:- (i) Ethyl Alcohol (ii) Benzyl Alcohol
(xii) Define Phenol. Give its two uses.

4. Attempt any six parts. **6 × 2 = 12**

(i) How will you distinguish chemically between Ethanol and Propanal?
(ii) How will you convert Ethanal into Lactic Acid?
(iii) How Amino Acids show basic character?
(iv) What happens when Ammonium Acetate is heated?
(v) What is Nylon – 6,6?
(vi) What is the difference between Fats and Oils?
(vii) What is Saponification Number? Give an example.
(viii) What happens when Ammonium Carbonate is heated?
(ix) State the raw materials used for manufacture of Cement.

SECTION-II

NOTE: - Attempt any three questions.

5.(a) Discuss essential features of Periods in Modern Periodic Table. 4
(b) Write four more important differences of Lithium from other Alkali metals. 4

6.(a) What do you mean by Corrosion? Explain electrochemical theory in detail. 4
(b) Discuss in detail the components of the Environment. 4

7.(a) What is Cracking? Explain types of Cracking. 4
(b) How will you carryout the following conversions:- 4
(i) $CH_4 \rightarrow CH_3CH_2COOH$ (ii) $CH_3 - CH_2 - CH_2 - Cl \rightarrow CH_3 - CH = CH_2$

8.(a) How will you prepare the following compounds starting from Ethyne? 4
(i) Acetaldehyde (ii) Acrylonitrile (iii) Oxalic Acid (iv) Ethane
(b) What is Aldol Condensation? Explain it with mechanism. 4

9.(a) Discuss the structure of Benzene on the basis of Atomic Orbital Treatment. 4
(b) Give chemical reactions to prepare Ethanol from Starch and Molasses by a Fermentation Process. 4

Paper Code Number: 4481

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

(1) Keeping in view the size of atoms, _____ order is correct one.

(A) $Mg > Sr$ (B) $Lu > Ce$ (C) $Ba > Mg$ (D) $C\ell > I$

(2) Chile saltpetre has the chemical formula:-

(A) KNO_3 (B) $Na_2B_4O_7$ (C) $NaNO_3$ (D) $Na_2CO_3.H_2O$

(3) Tincal is a mineral of:-

(A) $A\ell$ (B) Si (C) C (D) B

(4) Laughing gas is chemically:-

(A) N_2O (B) NO (C) NO_2 (D) N_2O_4

(5) _____ is the strongest acid.

(A) $HC\ell O$ (B) $HC\ell O_2$ (C) $HC\ell O_3$ (D) $HC\ell O_4$

(6) Group VI B of transition elements contains:-

(A) Cr, Mo, W (B) Zn, Cd, Hg (C) Fe, Ru, Os (D) $Mn, Te,$ Re

(7) ____ set of hybrid orbitals has planar triangular shape.

(A) sp^3 (B) sp^2 (C) sp (D) dsp^2

(8) Vinyl acetylene combines with $HC\ell$ to form:-

(A) Polyacetylene (B) Benzene (C) Chloroprene (D) Divinylacetylene

(9) During nitration of Benzene, the active nitrating agent is:-

(A) NO_3 (B) NO_2^+ (C) NO_2^- (D) HNO_3

(10) _____ is not a Nucleophile.

(A) H_2O (B) H_2S (C) NH_3 (D) BF_3

(11) Rectified spirit contains Alcohol about:-

(A) 80 % (B) 85 % (C) 90 % (D) 95 %

(12) Acetone reacts with HCN to form a cynohydrin. It is an example of:- (A) Nucleophilic addition (B) Electrophilic addition (C) Nucleophilic substitution (D) Electrophilic substitution

(13) Acetic acid is manufactured by:-

(A) Fermentation (B) Distillation (C) Ozonolysis (D) Esterification

(14) Acetamide is prepared by:- (A) Heating Ethyl Acetate (B) Heating Ammonium Acetate (C) Heating Methyl Cyanide (D) Heating Acetic Acid

(15) _____ polymers is an addition polymer.

(A) Polystyrene (B) Nylon – 6, 6 (C) Terylene (D) Epoxy resin

(16) Micro-nutrients are required in quantity ranging from:-

(A) 4 – 40 g (B) 6 – 200 g (C) 6 – 200 kg (D) 4 – 40 kg

(17) Ecosystem is a smaller unit of:-

(A) Atmosphere (B) Lithosphere (C) Biosphere (D) Hydrosphere

Paper Code Number: 4483

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

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(4) Keeping in view the size of atoms, _____ order is correct one.

(A) $Mg > Sr$ (B) $Lu > Ce$ (C) $Ba > Mg$ (D) $Cl > I$

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(7) Laughing gas is chemically:-

(A) N_2O (B) NO (C) NO_2 (D) N_2O_4

(8) _____ is the strongest acid.

(A) $HClO$ (B) $HClO_2$ (C) $HClO_3$ (D) $HClO_4$

(9) Group VI B of transition elements contains:-

(A) Cr, Mo, W (B) Zn, Cd, Hg (C) Fe, Ru, Os (D) $Mn, Te,$ Re

(10) ____ set of hybrid orbitals has planar triangular shape.

(A) sp^3 (B) sp^2 (C) sp (D) dsp^2

(11) Vinyl acetylene combines with HCl to form:-

(A) Polyacetylene (B) Benzene (C) Chloroprene (D) Divinylacetylene

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(A) NO_3 (B) NO_2^+ (C) NO_2^- (D) HNO_3

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(16) Acetic acid is manufactured by:-

(A) Fermentation (B) Distillation (C) Ozonolysis (D) Esterification

(17) Acetamide is prepared by:- (A) Heating Ethyl Acetate (B) Heating Ammonium Acetate (C) Heating Methyl Cyanide (D) Heating Acetic Acid

Paper Code
Number: 4485

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

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(A) Cr, Mo, W (B) Zn, Cd, Hg (C) Fe, Ru, Os (D) $Mn, Te,$ Re

(17) _____ set of hybrid orbitals has planar triangular shape.

(A) sp^3 (B) sp^2 (C) sp (D) dsp^2

Paper Code Number: 4487

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (NEW SCHEME) (SESSION 2015-2017)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

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(1) _____ is the strongest acid.

(A) $HClO$ (B) $HClO_2$ (C) $HClO_3$ (D) $HClO_4$

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(A) *Cr, Mo, W* (B) *Zn, Cd, Hg* (C) *Fe, Ru, Os* (D) *Mn, Te,* Re

(3) _____ set of hybrid orbitals has planar triangular shape.

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(16) Tincal is a mineral of:-

(A) Al (B) *Si* (C) *C* (D) *B*

(17) Laughing gas is chemically:-

(A) N_2O (B) NO (C) NO_2 (D) N_2O_4

2017 (S) Roll No: _____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 3.10 Hours **SUBJECTIVE** MAXIMUM MARKS: 83

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any Eight parts. **8 × 2 = 16**

(i) Why Fluoride Ion has bigger radius than the Fluorine Atom?

(ii) Why atomic radii decrease from left to right across a period?

(iii) What is a Dehydrating Agent? Give examples to show that Lime is a dehydrating agent.

(iv) What problems are faced by deficiency of Calcium in the plants?

(v) Write four uses of Sodium Silicate.

(vi) How will you convert Boric Acid into Borax and Vice Versa?

(vii) What is the action of cold and hot water on Phosphorous Pentoxide?

(viii) Why SO_3 is not dissolved in water in contact process?

(ix) Why Oxyacids of Chlorine are stronger than Oxyacids of Bromine?

(x) Why HF is weaker acid than HCl?

(xi) What is Corrosion? How does the water affect this process?

(xii) How Iron is Galvanized in Daily life?

3. Attempt any Eight parts. **8 × 2 = 16**

(i) Write a short note on Cracking of Hydrocarbons.

(ii) What are Homocyclic and Hetrocyclic Compounds? Give one example of each.

(iii) Write a short note on reactivity of Alkanes.

(iv) What is meant by Octane Number?

(v) Write a note on Friedel – Crafts Alkylation reaction.

(vi) How will you prepare the following compounds from benzene in two steps?
(a) m – chloronitrobenzene (b) p – chloronitrobenzene

(vii) Which is the best method of preparing Alkyl Halides? Give equation.

(viii) How Ethyl Magnesium Bromide is prepared in laboratory?

(ix) Why Ethyl Alcohol is liquid and Methyl Chloride a gas? Explain.

(x) How will you convert Methanol into Ethanol?

(xi) How is oil spillage affecting marine life?

(xii) What is Leachates?

4. Attempt any Six parts. **6 × 2 = 12**

(i) How Methanal can be prepared? Give reactions of:- (a) Laboratory method (b) Industrial method

(ii) How Ethanol can be prepared by Laboratory Method?
Why it is necessary to distilled off product immediately?

(iii) How carboxylic acid can be prepared from Alkanenitril?

(iv) Write two reactions of Carboxylic Acid in which Oxygen and Hydrogen bond is broken.

(v) How the Polymers are classified on the basis of process of Polymerization?
Give an example of each.

(vi) What is Sponification Number? Give its application.

P.T.O

(vii) Draw the parent nucleus of Steroids.

(viii) How urea is manufactured? Give its reaction involved.

(ix) Define Cement. Write its composition.

SECTION-II

NOTE: - Attempt any three questions. **8 ×3 = 24**

5.(a) Give two similarities and two differences of Hydrogen with group 1A elements. 4

(b) Describe commercial preparation of Sodium by Down's Cell. 4

6.(a) What happens when Orthoboric Acid reacts with

(i) C_2H_5OH (ii) $NaOH$ (iii) Na_2CO_3 (iv) H_2O 4

(b) Write important applications of Noble Gases. 4

7.(a) Write a detailed note on cracking of Hydrocarbons. 4

(b) What do you understand by the term β – elimination?
Explain briefly the E_1 mechanism with example. 4

8.(a) Discuss any four methods of preparation of Alkanes. 4

(b) What is Cannizzaro's reaction? Explain its reaction mechanism in detail. 4

9.(a) Write down four evidences which support Kekule's structure of Benzene. 4

(b) Explain with diagram industrial preparation of Methyl alcohol. 4

SECTION-III (PRACTICAL)

10. **Attempt any three parts.** **5 ×3 = 15**

(i) Write complete qualitative analysis for Zinc radical in a systematic manner. 5

(ii) Write complete qualitative analysis for Potassium radical in a systematic manner. 5

(iii) Write complete qualitative analysis for carbonate radical in a systematic manner. 5

(iv) How will you identify and confirm the carboxylic group in an organic compound? 5

(v) Write the material required, equation and procedure for the preparation of Copper Ammine Complex. 5

Paper Code
Number: 8481

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

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Q.No.1

(1) Statement _____ is incorrect.

(A) All metals are good conductors of heat (B) All metals are good conductors of electricity

(C) All metals form acidic oxides (D) All metals form positive ions

(2) _____ does not belong to Alkaline Earth metals.

(A) *Be* (B) *Ra* (C) *Ba* (D) *Rn*

(3) _____ metal is used in Thermit Process because of its activity.

(A) Iron (B) Copper (C) Aluminium (D) Zinc

(4) The highest ionization energy is possessed by:-

(A) Nitrogen (B) Phosphorus (C) Antimony (D) Bismith

(5) _____ is the weakest acid.

(A) *HF* (B) *HCℓ* (C) *HBr* (D) *HI*

(6) The total number of transition elements is:-

(A) 10 (B) 14 (C) 40 (D) 58

(7) The state of hybridization of Carbon atom in Methane is:-

(A) sp^3 (B) sp^2 (C) *sp* (D) dsp^2

(8) Formula of Chloroform is:-

(A) $CH_3Cℓ$ (B) $CH_2Cℓ_2$ (C) $CHCℓ_3$ (D) $CCℓ_4$

(9) _____ acid can be used as a catalyst in Friede–Crafts Reactions.

(A) $AℓCℓ_3$ (B) HNO_3 (C) $BaCℓ_2$ (D) $NaCℓ$

(10) _____ is not a nucleophile.

(A) H_2O (B) H_2S (C) BF_3 (D) NH_3

(11) _____ compound shows Hydrogen bonding.

(A) C_6H_6 (B) $C_2H_5Cℓ$ (C) CH_3-O-CH_3 (D) C_2H_5OH

(12) _____ will react with both the Aldehydes and Ketones.

(A) Grignards reagent (B) Tollen's reagent (C) Fehling's reagent (D) Benedict's reagent

(13) Acetic acid is manufactured by:-

(A) Distillation (B) Fermentation (C) Ozonolysis (D) Esterification

(14) _____ enzyme brings about the Hydrolysis of Fats.

(A) Urease (C) Maltase (C) Zymase (D) Lipase

(15) Phosphorus helps the growth of:-

(A) Root (B) Leave (C) Stem (D) Seed

(16) The main pollutant of leather tanneries in the waste water is due to salt of:-

(A) *Pb* (B) $Cr(VI)$ (C) *Cu* (D) $Cr(III)$

(17) Newspaper can be recycled again and again by _____ times.

(A) 2 (B) 3 (C) 4 (D) 5

Paper Code
Number: 8483

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

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(A) sp^3 (B) sp^2 (C) sp (D) dsp^2

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(A) $CH_3C\ell$ (B) $CH_2C\ell_2$ (C) $CHC\ell_3$ (D) $CC\ell_4$

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(17) Acetic acid is manufactured by:-

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Paper Code
Number: 8485

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

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(A) sp^3 (B) sp^2 (C) sp (D) dsp^2

(2) Formula of Chloroform is:-

(A) $CH_3C\ell$ (B) $CH_2C\ell_2$ (C) $CHC\ell_3$ (D) $CC\ell_4$

(3) _____ acid can be used as a catalyst in Friede–Crafts Reactions.

(A) $A\ell C\ell_3$ (B) HNO_3 (C) $BaC\ell_2$ (D) $NaC\ell$

(4) _____ is not a nucleophile.

(A) H_2O (B) H_2S (C) BF_3 (D) NH_3

(5) _____ compound shows Hydrogen bonding.

(A) C_6H_6 (B) $C_2H_5C\ell$ (C) CH_3-O-CH_3 (D) C_2H_5OH

(6) _____ will react with both the Aldehydes and Ketones.

(A) Grignards reagent (B) Tollen's reagent (C) Fehling's reagent (D) Benedict's reagent

(7) Acetic acid is manufactured by:-

(A) Distillation (B) Fermentation (C) Ozonolysis (D) Esterification

(8) _____ enzyme brings about the Hydrolysis of Fats.

(A) Urease (C) Maltase (C) Zymase (D) Lipase

(9) Phosphorus helps the growth of:-

(A) Root (B) Leave (C) Stem (D) Seed

(10) The main pollutant of leather tanneries in the waste water is due to salt of:-

(A) Pb (B) $Cr(VI)$ (C) Cu (D) $Cr(III)$

(11) Newspaper can be recycled again and again by _____ times.

(A) 2 (B) 3 (C) 4 (D) 5

(12) Statement _____ is incorrect.

(A) All metals are good conductors of heat (B) All metals are good conductors of electricity

(C) All metals form acidic oxides (D) All metals form positive ions

(13) _____ does not belong to Alkaline Earth metals.

(A) Be (B) Ra (C) Ba (D) Rn

(14) _____ metal is used in Thermit Process because of its activity.

(A) Iron (B) Copper (C) Aluminium (D) Zinc

(15) The highest ionization energy is possessed by:-

(A) Nitrogen (B) Phosphorus (C) Antimony (D) Bismith

(16) _____ is the weakest acid.

(A) HF (B) $HC\ell$ (C) HBr (D) HI

(17) The total number of transition elements is:-

(A) 10 (B) 14 (C) 40 (D) 58

Paper Code
Number: 8487

2017 (S) Roll No._____

INTERMEDIATE PART-II (12th CLASS)

CHEMISTRY PAPER-II (OLD SCHEME) (SESSION 2012-2014)

TIME ALLOWED: 20 Minutes **OBJECTIVE** MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.

Q.No.1

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BOARD OF INTERMEDIATE AND SECONDARY EDUCATION, MULTAN

OBJECTIVE KEY FOR INTER (PART-I / II) Annual Examination, 2017.

Name of Subject Chemistry (Old) II Session (15-17) Chemistry (New)

Group: 1st Session (12-14)

Q. Nos.	Paper Code 8481	Paper Code 8483	Paper Code 8485 (84)	Paper Code 8487
1.	C	D	A	A
2.	D	D	C	A
3.	C	B	A	D
4.	A	D	C	A
5.	A	C	D	C
6.	D	D	A	A
7.	A	C	B	C
8.	C	A	D	D
9.	A	A	D	A
10.	C	D	B	B
11.	D	A	D	D
12.	A	C	C	D
13.	B	A	D	B
14.	D	C	C	D
15.	D	D	A	C
16.	B	A	A	D
17.	D	B	D	C
18.				
19.				
20.				

Group: 2nd

Q. Nos.	Paper Code 4481	Paper Code 4483	Paper Code 4485	Paper Code 4487
1.	C	A	C	D
2.	C	B	B	A
3.	D	C	D	B
4.	A	C	D	C
5.	D	C	A	B
6.	A	D	A	D
7.	B	A	B	D
8.	C	D	A	A
9.	B	A	B	A
10.	D	B	C	B
11.	D	C	C	A
12.	A	B	C	B
13.	A	D	D	C
14.	B	D	A	C
15.	A	A	D	C
16.	B	A	A	D
17.	C	B	B	A
18.				
19.				
20.				

سرٹیفکیٹ بابت تصحیح سوالیہ پرچہ/مارکنگ Key

ہم نے مضمون کیمسٹری پرچہ Old/New گروپ ____ تعلیم ____ انٹرسالانہ امتحان 2017 کا سوالیہ پرچہ نیا/پرانا (Subjective & Objevtive) کو بنظر عمیق چیک کرلیا ہے یہ پرچہ سلیبس کے عین مطابق Set کیا گیا ہے۔اس سوالیہ پرچہ میں کسی قسم کی کوئی غلطی نہ ہے۔ہم نے سوالیہ پرچہ کا اردو اور انگریزی Version بھی چیک کرلیا ہے یہ Version آپس میں مطابقت رکھتے ہیں اور سلیبس (Syllabus) کے مطابق بھی ہیں۔نیز اس پرچہ کی Key کی بابت بھی تصدیق کی جاتی ہے کہ یہ بھی درست بنائی گئی ہے۔اس میں بھی کسی قسم کی کوئی غلطی نہ ہے۔

مزید یہ کہ ہم نے Key بنانے سے متعلق دفتر کی جانب سے تیارہ کردہ ہدایات وصول کر کے ان کا بغور مطالعہ کرلیا ہے اور ان کی روشنی میں Key بنائی ہے۔

PREPARED & CHECKED BY

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